

B7

(12) INTERNATIONAL APPLICATION PUBLISHED UNDER THE PATENT COOPERATION TREATY (PCT)

(19) World Intellectual Property Organization
International Bureau



(43) International Publication Date
19 October 2006 (19.10.2006)

PCT

(10) International Publication Number
WO 2006/108743 A1

(51) International Patent Classification:
C09G 1/02 (2006.01) **H01L 21/3105** (2006.01)
C09K 3/14 (2006.01)

(74) Common Representative: **DEGUSSA AG**; Intellectual Property Management Patente und Marken Standort Hanau, Postfach 13 45, 63403 Hanau (DE).

(21) International Application Number:
PCT/EP2006/060560

(81) Designated States (*unless otherwise indicated, for every kind of national protection available*): AE, AG, AL, AM, AT, AU, AZ, BA, BB, BG, BR, BW, BY, BZ, CA, CH, CN, CO, CR, CU, CZ, DE, DK, DM, DZ, EC, EE, EG, ES, FI, GB, GD, GE, GH, GM, HR, HU, ID, IL, IN, IS, JP, KE, KG, KM, KN, KP, KR, KZ, LC, LK, LR, LS, LT, LU, LV, LY, MA, MD, MG, MK, MN, MW, MX, MZ, NA, NG, NI, NO, NZ, OM, PG, PH, PL, PT, RO, RU, SC, SD, SE, SG, SK, SL, SM, SY, TJ, TM, TN, TR, TT, TZ, UA, UG, US, UZ, VC, VN, YU, ZA, ZM, ZW.

(22) International Filing Date: 8 March 2006 (08.03.2006)

(25) Filing Language: English

(26) Publication Language: English

(30) Priority Data:
10 2005 017 372.1 14 April 2005 (14.04.2005) DE

(71) Applicant (*for all designated States except US*): **DEGUSSA AG** [DE/DE]; Bennigsenplatz 1, 40474 Düsseldorf (DE).

(72) Inventors; and

(75) Inventors/Applicants (*for US only*): **KRÖLL, Michael** [DE/DE]; Schmidtgasse 18, 63571 Gelnhausen (DE). **HEBERER, Stefan** [DE/DE]; Wiesenbornstrasse 3, 63571 Gelnhausen (DE). **KATUSIC, Stipan** [DE/DE]; Berliner Ring 20, 65779 Kelkheim (DE). **KRÄMER, Michael** [DE/DE]; Dresdener Strasse 30, 61137 Schöneck-kilianstädten (DE). **LORTZ, Wolfgang** [DE/DE]; Feldstrasse 9, 63607 Wächtersbach (DE).

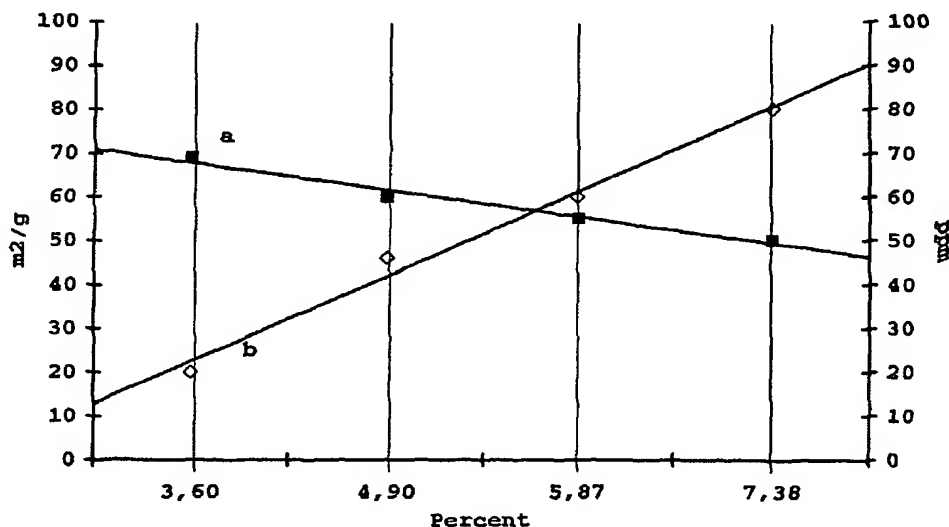
(84) Designated States (*unless otherwise indicated, for every kind of regional protection available*): ARIPO (BW, GH, GM, KE, LS, MW, MZ, NA, SD, SL, SZ, TZ, UG, ZM, ZW), Eurasian (AM, AZ, BY, KG, KZ, MD, RU, TJ, TM), European (AT, BE, BG, CH, CY, CZ, DE, DK, EE, ES, FI, FR, GB, GR, HU, IE, IS, IT, LT, LU, LV, MC, NL, PL, PT, RO, SE, SI, SK, TR), OAPI (BF, BJ, CF, CG, CI, CM, GA, GN, GQ, GW, ML, MR, NE, SN, TD, TG).

Published:

— with international search report

For two-letter codes and other abbreviations, refer to the "Guidance Notes on Codes and Abbreviations" appearing at the beginning of each regular issue of the PCT Gazette.

(54) Title: AQUEOUS CERIUM OXIDE DISPERSION



(57) Abstract: Aqueous cerium oxide dispersion Aqueous cerium oxide dispersion, containing 5 to 60% by weight cerium oxide. It can be used to polish SiO₂ in the semiconductor industry.

WO 2006/108743 A1

Aqueous cerium oxide dispersion

The invention relates to a cerium oxide dispersion, to a process for producing it and to its use.

Cerium oxide dispersions are used in the semiconductor industry to polish SiO_2 (CMP). Cerium oxide dispersions are known from EP 1 596 940. These cerium oxide dispersions have particles with a mean aggregate diameter d_{50} of less than 200 nm.

The known cerium oxide dispersion disadvantageously includes at least one dispersant.

In addition, the known cerium oxide dispersion includes further additives which are used for STI CMP and are intended to shift the selectivity during the polishing of SiO_2 and Si_3N_4 in favour of the SiO_2 .

The known cerium oxide dispersion also has the drawback that it has SiO_2 removal rates of < 600 nm/min and is additionally unstable, since the cerium oxide forms a sediment (cf. DE 101 62 174 A1, US 6,312,486).

The object therefore exists of producing a cerium oxide dispersion which does not have these drawbacks.

The subject matter of the application is an aqueous cerium oxide dispersion, containing 5 to 60% by weight cerium oxide, preferably 20 to 50% by weight cerium oxide, in particular 35 to 45% by weight cerium oxide.

This dispersion can be produced by means of wet jet mill or ultrasound equipment.

The cerium oxide dispersion according to the invention does not contain any dispersion additives or any additives for increasing the selectivity in STI CMP. It may have a pH of from 3 to 8. Since the dispersion according to the invention does not contain any dispersion additives, the stabilizing can only be effected by means of the pH.

It may have a low viscosity, in particular a viscosity of < 30 mPas, preferably < 20 mPas.

The particles of the dispersion according to the invention may have a particle size distribution, determined using Horiba LA 910, as follows:

d50 < 120 nm, preferably d50 < 100 nm, in particular
d50 < 90 nm
d99 < 200 nm, preferably d99 < 150 nm, in particular
d99 < 130 nm.

The dispersion according to the invention has an excellent stability. The particle size distribution does not change with prolonged storage of the dispersion.

The cerium oxide dispersion according to the invention can be produced using known cerium oxide powders.

By way of example, it is possible to produce a cerium oxide powder as described in EP 1 506 940 A1.

Furthermore, it is possible to produce a cerium oxide powder in accordance with German Patent Application DE 10 2005 005 344.0. This application was filed with the German Patents and Trademark Office on 5 February 2005.

Its wording is as follows, with regard to the cerium oxide powder:

It is an object of the invention to provide a process for producing metal oxide powders which allows a uniform powder with a high BET surface area and a narrow particle size distribution to be achieved and which is suitable for producing quantities in the kg/h range.

The subject matter of the invention is a process for producing a metal oxide powder with a BET surface area of at least $20 \text{ m}^2/\text{g}$ by reacting an aerosol with oxygen in a reaction space at a reaction temperature of more than 700°C , followed by separation of the powder obtained from the gaseous substances, wherein

- the aerosol is obtained by spraying at least one starting material, either itself in liquid form or in solution, and at least one atomization gas by means of a multi-fluid nozzle,
- the volume-based, mean drop diameter D_{30} of the aerosol is 30 to 100 μm , and
- the number of aerosol drops which are larger than 100 μm is up to 10% based on the total number of drops.

Metal oxide powders in the sense of the invention also encompass mixed metal oxide powders and doped metal oxide powders.

A mixed metal oxide powder is to be understood as meaning a powder involving intimate mixing of mixed oxide components at the level of primary particles or aggregates. In this case, the primary particles may have oxygen-bridged metal components in the form of $M_1\text{-O-M}_2$ bonds. In addition, it is also possible for regions of individual oxides $M_1\text{O}$, $M_2\text{O}$, $M_3\text{O}$, ... to be present within the primary particles.

A doped metal oxide powder is to be understood as meaning a powder in which the doping component is located predominantly or exclusively at one lattice site in the metal oxide lattice. The doping component may be in metallic or oxidic form. An example of a doped metal oxide powder is indium tin oxide, where tin atoms occupy sites in the lattice of the indium oxide.

The volume-based, mean drop diameter D_{30} is calculated as follows:

$$D_{30} = \sqrt[3]{\frac{1}{N} \sum_{i=1}^N D_i^3}$$

A starting material is to be understood as meaning a metal compound which is converted into a metal oxide under the reaction conditions. In the case of doped metal oxides, it may be possible for the starting material of the doping component to be converted into a metallic component.

In the context of the present process according to the invention, it had been discovered that it is possible to produce metal oxide powders with a high specific surface area by virtue of the volume-based, mean drop diameter D_{30} being from 30 to 100 μm and at the same time up to 10% of the absolute number of drops being larger than 100 μm . It is in this way possible to increase the throughput of solution compared to the prior art without having to accept a significant reduction in the BET surface areas of the powders. The BET surface area of the powders obtained by the process according to the invention is at least 20 m^2/g , preferably 20 to 200 m^2/g .

The absolute drop size is determined using the principle of dual phase Doppler Anemometry using a 5W Argon Ion continuous wave laser.

In a preferred embodiment, the number of drops, based on the total number of drops, which are larger than 100 μm may be from 3% to 8%.

Furthermore, it may be advantageous if the percentage of drops which are larger than 250 μm is no more than 10%, based on the number of drops > 100 μm .

In particular, it may be advantageous to use an embodiment in which the following dependent relationship between the volume-based, mean drop diameter D_{30} and the spray width of the aerosol applies:

Spray width [mm]	D_{30} [μm]
0	10 to 30
± 20	20 to 40
± 40	30 to 60
± 60	50 to 80
± 80	80 to 120.

The throughput of a solution containing a starting material may preferably be from 1.5 to 2000 kg/h and particularly preferably from 100 to 500 kg/h.

In this case, the starting material content in the solution may be 2 to 60, preferably 5 to 40% by weight.

The metal component of the starting material may preferably be Al, Ce, Fe, Mg, In, Ti, Si, Sn, Y, Zn and/or Zr. In the context of the present invention, silicon dioxide is to be regarded as a metal oxide. Starting materials comprising Al, Ce or Zn may be preferred.

If starting materials with different metal components are used, mixed metal oxide powders are obtained. In this case, the proportion of one specific starting material and therefore the proportion of the corresponding metal oxide

component in the mixed metal oxide powder is not subject to any particular restrictions.

The starting materials may be organometallic and/or inorganic in form, in which context organometallic compounds may be preferred. Examples of inorganic starting materials may in particular be metal chlorides and metal nitrates. Organometallic compounds used may in particular be metal alkoxides and/or metal carboxylates. Alkoxides used may preferably be ethoxides, n-propoxides, isopropoxides, n-butoxides and/or tert-butoxides. The carboxylates used may be those of acetic acid, propionic acid, butanoic acid, hexanoic acid, oxalic acid, malonic acid, succinic acid, glutaric acid, adipic acid, octanoic acid, 2-ethylhexanoic acid, valeric acid, capric acid and/or lauric acid. 2-ethylhexanoates and/or laurates may particularly advantageously be used.

The inorganic starting compounds may preferably be in the form of a solution in water, and the organometallic starting compounds may be in the form of a solution in organic solvents.

Organic solvents or constituents of organic solvent mixtures used may preferably be alcohols, such as methanol, ethanol, n-propanol, iso-propanol, n-butanol or tert-butanol, diols, such as ethanediol, pentanediol, 2-methyl-2,4-pentanediol, dialkyl ethers, such as diethyl ether, tert-butyl-methyl ether or tetrahydrofuran, C₁-C₁₂-carboxylic acids, such as for example acetic acid, propionic acid, butanoic acid, hexanoic acid, oxalic acid, malonic acid, succinic acid, glutaric acid, adipic acid, octanoic acid, 2-ethylhexanoic acid, valeric acid, capric acid, lauric acid. It is also possible to use ethyl acetate, benzene, toluene, naphtha and/or benzene. Solutions which contain C₂-C₁₂-carboxylic acids, in particular 2-ethylhexanoic acid and/or lauric acid, may preferably be used.

It is preferable for the C₂-C₁₂-carboxylic acid content in the solution to be less than 60% by weight, particularly preferably less than 40% by weight, based on the total quantity of solution.

In a particularly preferred embodiment, the solutions of the starting materials simultaneously contain a carboxylate and the carboxylic acid on which it is based and/or an alkoxide and the alcohol on which it is based. The starting materials used may in particular be the 2-ethylhexanoates in a solvent mixture which contains 2-ethylhexanoic acid.

The atomization gas used in the process according to the invention may be a reactive gas, such as air, oxygen-enriched air and/or an inert gas such as nitrogen. The atomization gas used is generally air.

According to the process of the invention, the quantity of atomization gas is such that the ratio throughput of the solution of the starting material/quantity of atomization gas is preferably 2 to 25 kg/m³ (s.t.p.), particularly preferably 5 to 10 kg/m³ (s.t.p.).

Suitable multi-fluid nozzles for the process according to the invention are in particular three-fluid nozzles or four-fluid nozzles.

If three-fluid nozzles or four-fluid nozzles are used, it is possible for two or three, respectively, separate solutions, which contain

- identical or different starting materials,
 - in identical or different solvents,
 - in identical or different concentration,
- to be atomized in addition to the atomization gas.

By way of example, two solutions with different concentrations of a starting material and the same solvent

or solvent mixture can be atomized simultaneously in this way. As a result, aerosol drops of different sizes are obtained.

Furthermore, by way of example, the atomization gas can be supplied via two nozzles or it is possible to use different atomization gases, for example air and steam.

Separate solutions of different starting materials can be used to produce mixed oxide powders.

The reaction temperature of more than 700°C required in the process according to the invention can preferably be obtained by means of a flame generated by reaction of a hydrogen-containing fuel gas with (primary) air, if appropriate enriched with oxygen. Suitable fuel gases may be hydrogen, methane, ethane, propane, butane and/or natural gas, with hydrogen being particularly preferred. The reaction temperature is defined as the temperature which is established 0.5 m below the flame.

Furthermore, it may be advantageous if secondary air is additionally introduced into the reaction space. The quantity of secondary air will generally be set in such a way that the ratio of secondary air to primary air is from 0.1 to 10.

It is particularly advantageous if λ is ≥ 1.5 , with λ being calculated from the quotient of the sum of the oxygen content of the air used (primary air, secondary air and atomization air) divided by the sum of the starting materials and of the hydrogen-containing fuel gas, in each case in mol/h. Very particularly preferably, the condition $2 < \lambda < 5$ may apply.

A cooling process generally precedes the separation of the powder out of the reaction mixture. This cooling process can be implemented directly, for example by means of a

quench gas, or indirectly, for example by means of external cooling.

A further subject of the invention is a metal oxide powder obtainable by the process according to the invention. The metal oxide powder may contain impurities caused by the starting material and/or the process. The purity of the calcined metal oxide powder is at least 98% by weight, generally at least 99% by weight. A content of at least 99.8% by weight may be particularly preferred.

In general, the metal oxide powder is predominantly or exclusively in the form of aggregates of primary particles, in which case the aggregates do not have cenospheric structures. In the context of the invention, a cenospheric structure is to be understood as meaning a structure which has a size of from 0.1 to 20 μm and is approximately in the shape of a hollow sphere with a wall thickness of from 0.1 to 2 μm . The term predominantly is to be understood as meaning that a TEM image reveals individual unaggregated particles forming a proportion of no more than 10%.

The metal oxide powder may preferably have a BET surface area of from 30 to 200 m^2/g .

The coarse fraction $> 45 \mu\text{m}$ in the metal oxide powder according to the invention is preferably less than 100 ppm, particularly preferably less than 50 ppm.

The metal oxide powder according to the invention preferably has a carbon content of less than 0.15% by weight and a chloride, sodium and potassium content of less than 300 ppm.

The metal oxide powder according to the invention may preferably be a cerium oxide powder with a BET surface area of from 30 to 90 m^2/g .

The cerium oxide powder according to the invention, exposed to air and temperatures of 900°C for a period of two hours, may have a BET surface area of up to 35 m²/g.

The mean primary particle diameter of the cerium oxide powder may preferably be 5 to 20 nm, particularly preferably 8 to 14 nm.

The mean aggregate diameter of the cerium oxide powder may be 20 to 100 nm, particularly preferably 30 to 70 nm.

A further subject of the invention is the use of the metal oxide powder according to the invention for the production of dispersions, for the polishing of glass and metal surfaces, as a catalyst and as a catalyst support.

Examples

The specific surface area is determined in accordance with DIN 66131.

The TEM images are obtained using a Hitachi TEM appliance, type H-75000-2. The primary particle diameter and aggregate diameter of in each case approx. 2000 aggregates are evaluated by means of CCD-camera of the TEM appliance and subsequent image analysis. For this purpose, the powders are dispersed in isopropanol/water (1 min, ultrasonic processor UP 100 H, Dr. Hielscher GmbH, 100 W).

The level of particles larger than 45 μm is determined based on DIN EN ISO 787, part 18.

Starting materials

Solution A: 42% by weight cerium(III)ethylhexanoate, 25% by weight 2-ethylhexanoic acid, 4% by weight 2-methyl-2,4-pentanediol, naphtha 29% by weight.

Solution B: 30% by weight cerium(III) acetate hydrate, 50% by weight acetic acid, 20% by weight lauric acid.

Solution C: 24.4 zirconium(III) ethylhexanoate, 0.30% by weight hafnium(III) ethylhexanoate, 39.6% by weight 2-ethylhexanoic acid, 3.5% by weight 2-(2-butoxyethoxy)ethanol, 32.2% by weight white spirit

Example 1:

An aerosol is generated from the part-streams I: 200 kg/h of solution A, II: 50 kg/h of solution A und III: 17.3 m³/h (s.t.p.) of atomization air by means of a three-fluid nozzle (Schlick, Model 0/4 S41), and this aerosol is atomized into a reaction space, where a hydrogen/oxygen flame comprising hydrogen (40 m³/h (s.t.p.)) and primary air (1800 m³/h (s.t.p.)), in which the aerosol is reacted,

is burning. In addition, secondary air ($3200 \text{ m}^3/\text{h}$ (s.t.p.)) is introduced into the reaction space.

After cooling, the cerium oxide powder is separated from gaseous substances at a filter. The residence time of the reaction mixture in the reaction space is 0.9 s. The temperature 0.5 m below the flame is 1100°C .

Examples 2 to 4 are carried out analogously to Example 1. Part-streams I and II remain identical, whereas the throughput and pressure of part-stream III are increased.

Example 5 is carried out analogously to Example 1, except that solution B is used instead of solution A for part-stream II.

Example 6 is carried out analogously to Example 1, except that solution C is used instead of solution A.

Example 7 is carried out analogously to Example 1, but with part-stream I consisting of solution A and part-stream II consisting of solution C.

Table 1 shows the parameters which are of significance to the aerosol generation, and Table 2 shows the parameters which are of significance for the flame.

Table 3 shows the analytical values of the powders obtained.

Table 4 shows the drop distribution in percent of all the drops $>100 \text{ }\mu\text{m}$ from Examples 1 to 4.

Figure 1 shows the dependency of the BET surface area(a) in m^2/g and of the particles $> 45 \text{ }\mu\text{m}$ (b) on the percentage of drops $> 100 \text{ }\mu\text{m}$. Figure 1 shows the relevance of the proportion of drops $> 100 \text{ }\mu\text{m}$ to the BET surface area and the coarse fraction $> 45 \text{ }\mu\text{m}$. The process according to the

invention allows the production of large quantities of metal oxide powder with a high BET surface area and a low coarse fraction, it being possible for the BET surface area and coarse fraction to be set by means of the proportion of drops $> 100 \mu\text{m}$.

Figure 2 shows the D_{30} drop diameter [μm] of the powders from Examples 1 to 4 in μm as a function of the spray width [mm]. D_{30} values of up to $180 \mu\text{m}$ are obtained at the edges. Nevertheless, the process according to the invention allows the production of fine-particle metal oxide powders.

Table 1: Aerosol generation

Example	1	2	3	4	5	6	7
Part-stream I							
Solution	A	A	A	A	A	C	A
Flow rate [kg/h]	50.0	50.0	50.0	50.0	50.0	50.0	70
Part-stream II							
Solution	A	A	A	A	B	C	C
Flow rate [kg/h]	200.0	200.0	200.0	200.	200.0	200.0	170
Part-stream III							
Atomization air							
[m ³ /h (s.t.p)]	17.3	28.0	35.8	42.9	17.3	17.3	17.3
Pressure [bar superatmospheric]							
Part-stream I	0.4	0.4	0.4	0.4	0.4	0.4	0.4
Part-stream II	4.3	4.3	4.3	4.3	4.3	4.3	4.3
Part-stream III	0.5	1.0	1.5	2.0	0.5	0.5	0.5
Drops							
Total number	598864	539612	414543	381120	598812	598235	539009
Diameter							
< 100 μm	92.62	94.13	95.10	98.94	92.81	92.15	91.60
> 100 μm	7.38	5.87	4.90	3.60	7.29	7.85	8.40

Table 2: Flame parameter

Example	1	2	3	4	5	6	7
Hydrogen [m ³ /h (s.t.p.)]	40	40	40	40	40	40	45
Primary air [m ³ /h (s.t.p.)]	1800	1800	1800	1800	1800	1800	1800
Secondary air [m ³ /h (s.t.p.)]	3200	3200	3200	3200	3200	3200	3300
Atomization air [m ³ /h (s.t.p.)]	17.3	28.0	35.8	42.9	17.3	17.3	17.3
lambda	2.63	2.64	2.64	2.64	2.63	2.63	2.65
Residence time [s]	0.90	0.90	1.12	1.12	0.90	0.90	0.85
Temperature ^{a)} [°C]	1100	1020	1050	1050	1100	1100	1130

a) Temperature = 0.5 m beneath the flame

Table 3: Analytical values for the powders obtained

Example	1	2	3	4	5	6	7
BET surface area [m ² /g]	50	55	60	69	42	45	70
Fractions >45 µm [ppm]	60	80	46	20	82	90	38
C content [% by weight]	0.08	0.08	0.09	0.09	0.11	0.13	0.07
Cl-Content ^{a)} [ppm]	225	251	161	242	261	184	192
Na-Content ^{a)} [ppm]	176	201	160	189	212	161	180
K-Content ^{a)} [ppm]	112	156	145	126	154	128	167
Mean primary particle diameter ^{b)} [nm]	9.3	11.3	9.0	8.5	-	13.0	9.6
Mean aggregate diameter ^{b)} [nm]	47.5	48.5	47.3	47.0	-	66.0	47.1
Mean aggregate surface area ^{b)} [nm ²]	2410	2738	2312	2280	-	5228	4288

a) According to ICP-Measurement; b) Determined by means of image analysis;

Table 4: Drop distribution in % of all drops >100 μm

Example1	1	2	3	4
100 - 150 μm	7.14	11.17	18.00	25.39
151 - 200 μm	42.09	48.67	56.42	37.23
201 - 250 μm	45.64	33.25	23.33	33.34
> 250 μm	5.13	6.91	2.25	4.04

A further subject of the invention is the use of the cerium oxide dispersion according to the invention for polishing glass, SiO_2 or SiO_2 on Si_3N_4 in the semiconductor industry.

During STI-CMP, the end point of the polishing of SiO_2 can be determined by means of the drop in the motor current when the nitride layer is reached.

A cerium oxide dispersion according to the invention can be used to polish SiO_2 surfaces with material removal rates of > 600 nm/min, preferably up to 1200 nm/min, in particular 1000 nm/min at a small pV value of 60000 N/ms. For this purpose, a slurry is produced containing 0.1 to 10% cerium oxide, preferably 0.2 to 5% and particularly preferably 0.5 - 1.5% cerium oxide. This dispersion can be obtained by dilution. Polishing is carried out at pHs of < 8, preferably 5 - 8, particularly preferably 7 - 8.

Moreover, it was surprisingly established that the motor currents measured during the polishing of SiO_2 and Si_3N_4 differ from one another if the cerium oxide dispersion according to the invention is used. This is unusual and can be exploited in what is known as STI (Shallow Trench Isolation). Here, a structure comprising SiO_2 and Si_3N_4 is produced on the wafer.

The objective is to polish SiO_2 as selectively as possible and to stop the polishing operation when the nitride is reached. For this purpose, additives are added to the slurry so as to increase this selectivity by protecting the nitride (protection against over-polishing) The end point determination is carried out by the determination of NH_3 .

Surprisingly, these additives are not required with the dispersion according to the invention.

Since the motor currents measured during polishing vary considerably, the drop in the motor current can be used for the end point determination (if the motor current drops, the polishing of the oxide is complete and the nitride has been reached → intrinsic protection against over-polishing even without additives).

Figure 3 shows the particle size distribution (Horiba LA 910) of a 5% strength dispersion in water at a pH of 3.9 and without additives.

Figure 4 shows the viscosity of a 40% strength dispersion in water.

Figure 5 shows the motor currents measured for SiO_2 polishing and Si_3N_4 polishing using a 0.5% strength cerium oxide dispersion with a surface area of $50 \text{ m}^2/\text{g}$ in water, at a pH of 7.6.

Figure 6 shows SiO_2 polishing tests using a 0.5% strength cerium oxide dispersion, at a pH of 7.6, the cerium oxide having a BET surface area of $50 \text{ m}^2/\text{g}$.

Patent Claims

1. Aqueous cerium oxide dispersion, containing 5 to 60% by weight cerium oxide.
2. Aqueous cerium oxide dispersion according to Claim 1, characterized in that it does not contain any dispersants or any additives for increasing the selectivity in STI CMP.
3. Aqueous cerium oxide dispersion according to Claim 1 or 2, characterized in that it has a pH of from 3 to 8.
4. Aqueous cerium oxide dispersion according to Claim 1, characterized in that it has a viscosity of less than 30 mPas.
5. Aqueous cerium oxide dispersion according to Claim 1, characterized in that it includes aggregated particles with a particle size distribution of
d50 < 120 nm
d99 < 200 nm.
6. Process for producing the cerium oxide dispersion according to Claim 1, characterized in that a cerium oxide powder is dispersed in water.
7. Use of the cerium oxide dispersion according to Claim 1 for polishing SiO₂ or Si₃N₄ in the semiconductor industry.
8. Use of the cerium oxide dispersion according to Claim 1 for polishing SiO₂ on Si₃N₄ in the semiconductor industry.
9. Process for polishing SiO₂ on Si₃N₄ using the cerium oxide dispersion according to Claim 1, characterized in that the end point of the polishing of the SiO₂ is determined by means of the drop in the motor current.

1/4

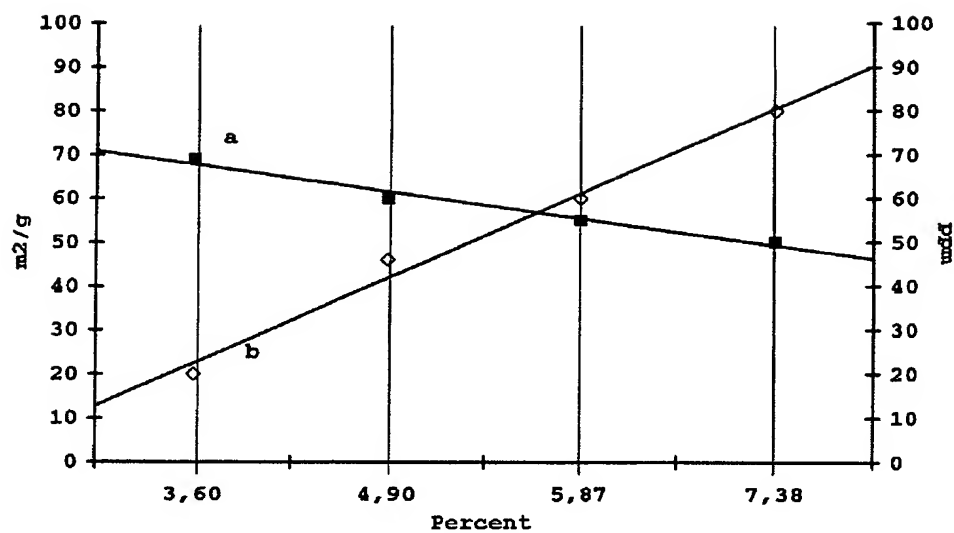


Figure 1

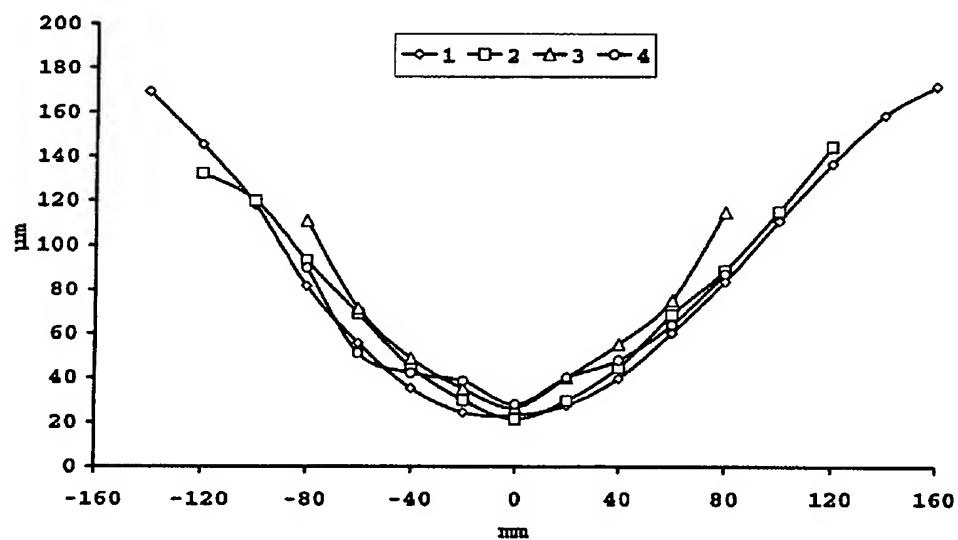


Figure 2

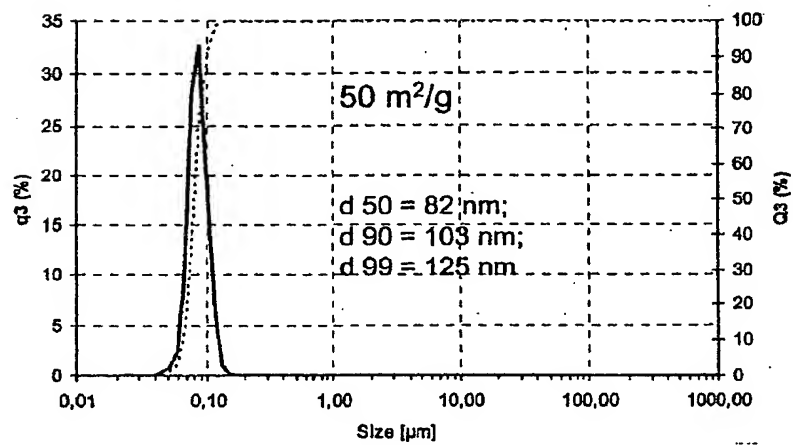


Figure 3

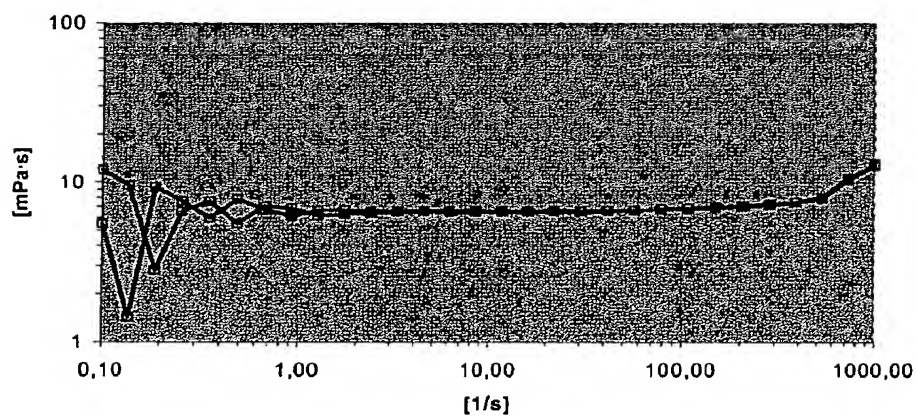


Figure 4

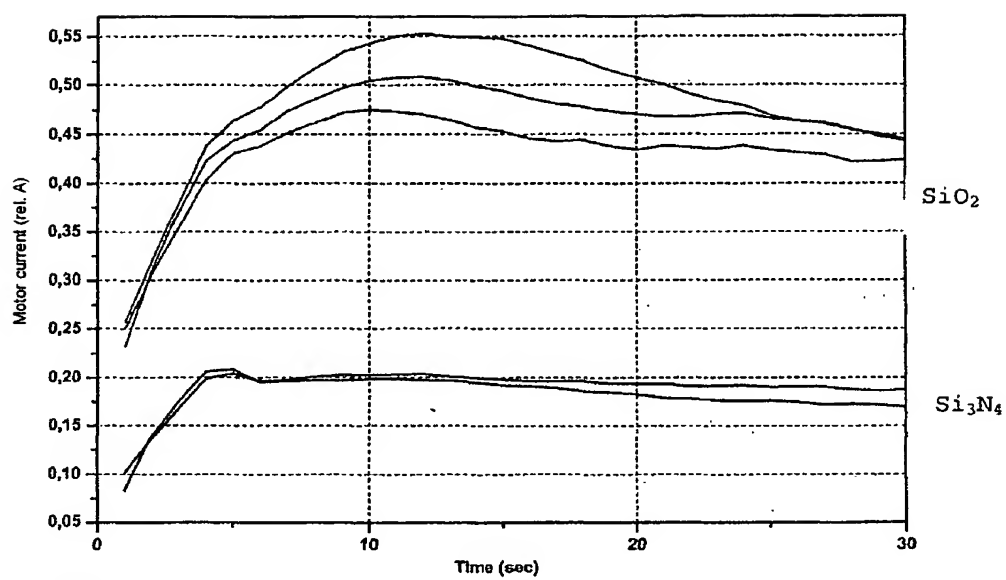


Figure 5

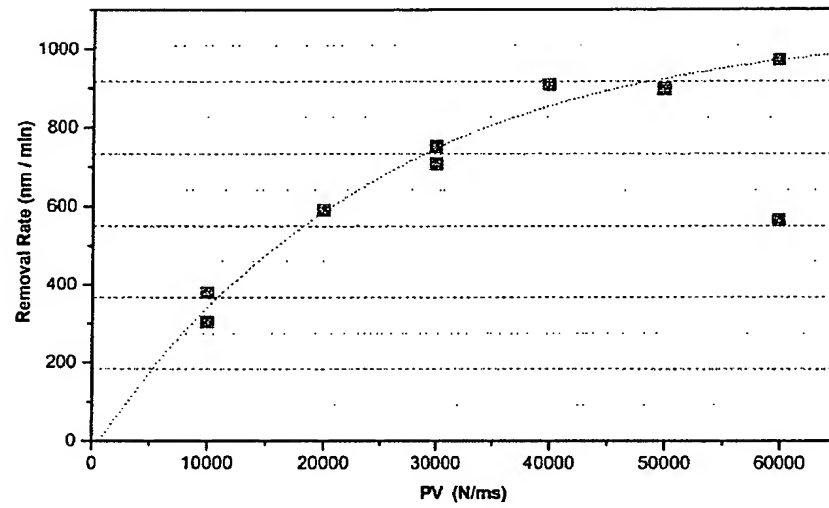


Figure 6

INTERNATIONAL SEARCH REPORT

International application No
PCT/EP2006/060560

A. CLASSIFICATION OF SUBJECT MATTER
INV. C09G1/02 C09K3/14 H01L21/3105

According to International Patent Classification (IPC) or to both national classification and IPC

B. FIELDS SEARCHED

Minimum documentation searched (classification system followed by classification symbols)
C09G C09K H01L

Documentation searched other than minimum documentation to the extent that such documents are included in the fields searched

Electronic data base consulted during the International search (name of data base and, where practical, search terms used)

EPO-Internal

C. DOCUMENTS CONSIDERED TO BE RELEVANT

Category*	Citation of document, with indication, where appropriate, of the relevant passages	Relevant to claim No.
X	EP 1 493 789 A (JSR CORPORATION) 5 January 2005 (2005-01-05)	1-8
Y	paragraphs [0001], [0005] - [0020], [0044] - [0066]; claims; examples	9
Y	US 2001/019895 A1 (ROBINSON KARL M ET AL) 6 September 2001 (2001-09-06) paragraphs [0001], [0011] - [0044]; claims	9
X	EP 1 274 123 A (SHOWA DENKO K.K.) 8 January 2003 (2003-01-08)	1-8
Y	paragraphs [0001], [0013] - [0035], [0045] - [0053]; claims; examples	9
	----- -/--	

☒ Further documents are listed in the continuation of Box C.

☒ See patent family annex.

* Special categories of cited documents:

- "A" document defining the general state of the art which is not considered to be of particular relevance
- "E" earlier document but published on or after the international filing date
- "L" document which may throw doubts on priority claim(s) or which is cited to establish the publication date of another citation or other special reason (as specified)
- "O" document referring to an oral disclosure, use, exhibition or other means
- "P" document published prior to the international filing date but later than the priority date claimed

- "T" later document published after the international filing date or priority date and not in conflict with the application but cited to understand the principle or theory underlying the invention
- "X" document of particular relevance; the claimed invention cannot be considered novel or cannot be considered to involve an inventive step when the document is taken alone
- "Y" document of particular relevance; the claimed invention cannot be considered to involve an inventive step when the document is combined with one or more other such documents, such combination being obvious to a person skilled in the art
- "&" document member of the same patent family

Date of the actual completion of the international search

22 June 2006

Date of mailing of the international search report

30/06/2006

Name and mailing address of the ISA/

European Patent Office, P.B. 5818 Patentlaan 2
NL - 2280 HV Rijswijk
Tel. (+31-70) 340-2040, Tx. 31 651 epo nl,
Fax: (+31-70) 340-3016

Authorized officer

Otegui Rebollo, J

INTERNATIONAL SEARCH REPORT

International application No

PCT/EP2006/060560

C(Continuation). DOCUMENTS CONSIDERED TO BE RELEVANT

Category*	Citation of document, with indication, where appropriate, of the relevant passages	Relevant to claim No.
Y	US 5 036 015 A (SANDHU ET AL) 30 July 1991 (1991-07-30) column 1, lines 7-10 column 2, line 20 - column 5, line 66; claims; figures	9
X	EP 1 130 630 A (SHOWA DENKO KABUSHIKI KAISHA) 5 September 2001 (2001-09-05)	1-8
Y	paragraphs [0001] - [0011], [0016] - [0020], [0024] - [0030]; claims; examples	9
X	US 2005/074610 A1 (KROLL MICHAEL ET AL) 7 April 2005 (2005-04-07) paragraphs [0001], [0008] - [0034]; claims	1-8
X	EP 1 234 801 A (NISSAN CHEMICAL INDUSTRIES, LTD) 28 August 2002 (2002-08-28) paragraphs [0001], [0004] - [0039]; claims; examples	1-8
X	EP 1 043 379 A (HITACHI CHEMICAL COMPANY, LTD) 11 October 2000 (2000-10-11) paragraphs [0001], [0004] - [0020], [0025] - [0051]; claims; examples	1-8
X	EP 1 219 568 A (NISSAN CHEMICAL INDUSTRIES, LTD) 3 July 2002 (2002-07-03) paragraphs [0001], [0007] - [0028], [0046] - [0053]; claims; examples	1-8
X	EP 1 179 578 A (PRAXAIR S.T. TECHNOLOGY, INC) 13 February 2002 (2002-02-13) paragraphs [0007], [8.10] - [0018]; claims; examples	1-6

INTERNATIONAL SEARCH REPORT

Information on patent family members

International application No

PCT/EP2006/060560

Patent document cited in search report		Publication date	Patent family member(s)	Publication date
EP 1493789	A	05-01-2005	CN 1576345 A JP 2005026604 A US 2005001199 A1	09-02-2005 27-01-2005 06-01-2005
US 2001019895	A1	06-09-2001	US 6191037 B1 US 2001010972 A1 US 2001009811 A1	20-02-2001 02-08-2001 26-07-2001
EP 1274123	A	08-01-2003	AU 4875601 A WO 0180296 A1 TW 586157 B	30-10-2001 25-10-2001 01-05-2004
US 5036015	A	30-07-1991	NONE	
EP 1130630	A	05-09-2001	WO 0008678 A1 JP 3560484 B2 JP 2000114211 A TW 419737 B	17-02-2000 02-09-2004 21-04-2000 21-01-2001
US 2005074610	A1	07-04-2005	DE 10342826 B3	12-05-2005
EP 1234801	A	28-08-2002	CN 1371867 A US 2003007920 A1	02-10-2002 09-01-2003
EP 1043379	A	11-10-2000	AU 1683899 A CA 2315057 A1 CN 1480503 A CN 1128195 C WO 9931195 A1 JP 3727241 B2 TW 577913 B US 6343976 B1	05-07-1999 24-06-1999 10-03-2004 19-11-2003 24-06-1999 14-12-2005 01-03-2004 05-02-2002
EP 1219568	A	03-07-2002	CN 1361064 A US 2002086618 A1	31-07-2002 04-07-2002
EP 1179578	A	13-02-2002	CN 1329118 A JP 2002028850 A SG 94358 A1 TW 550286 B US 6454821 B1	02-01-2002 29-01-2002 18-02-2003 01-09-2003 24-09-2002